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TABLE OF CONTENTS AND INDEX TO VOLUME TWELVE

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TABLE OF CONTENTS

No. 1 - FEBRUARY

Observations on citrus diseases in the Philippines, <i>J.M. Wallace and A. Martinez</i>	1
Pear moria in Italy, <i>E. Refatti</i>	6
<i>Alternaria sesami</i> , a serious seed-borne pathogen of worldwide distribution, <i>E.E. Leppik and G. Sowell Jr.</i>	13
Recommended analytical methods for pesticides	
22. Chlorfenson and chlorfenson formulations	17
<i>Collaborative Pesticides Analytical Committee</i>	
Outbreaks and new records	
Rhodesia and Nyasaland	22
Plant quarantine announcements	
Australia, Ivory Coast, Luxembourg	23
News and notes	
	24

No. 2 - APRIL

Highlights of insect conditions in the United States in 1963, <i>William L. Seal</i>	25
World citrus problems	
III. Syria, <i>L.C. Knorr and J.R. Vaughn</i>	37
Notes on some parasites, hyperparasites and predators of coconut pests in Malaya, <i>R.J.A. Lever</i>	42
Recommended analytical methods for pesticides	
23. Fenson and fenson formulations	44
<i>Collaborative Pesticides Analytical Committee</i>	
Plant quarantine announcements	
United Kingdom (Great Britain)	46
News and notes	
	48

No. 3 - JUNE

Plant-parasitic nematodes in the United Arab Republic, <i>A.C. Tarjan</i>	49
Citrus virus diseases, <i>S. Moreira</i>	57
Present status of research on lethal yellowing disease of coconut palm in Jamaica, <i>Walter Carter</i>	67
Recommended analytical methods for pesticides	
24. Demeton-ethyl, demeton-methyl	70
<i>Collaborative Pesticides Analytical Committee</i>	
Plant quarantine announcements	
United Kingdom, Upper Volta	72

No. 4 - AUGUST

Studies on the spread of lethal yellowing disease of the coconut palm, <i>Walter Carter and Joseph R.R. Suah</i>	73
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Plant diseases in Samoa, <i>A. Johnston</i>	79
<i>Parlatoria cinerea</i> , a pest of citrus in Israel, <i>Uri Gerson</i>	82
Recommended analytical methods for pesticides	
25. Demeton-S-methyl	
<i>Collaborative Pesticides Analytical Committee</i>	86
26. 2,2-dichloropropionic acid and sodium 2,2-dichloropropionate	
<i>Collaborative Pesticides Analytical Committee</i>	88
Outbreaks and new records	
Brazil, Senegal, United States	92
Plant quarantine announcements	
Honduras	94

No. 5 - OCTOBER

Citrus false exanthema induced by feeding of a myrid, <i>Gerd Walter Müller and A.S. Costa</i>	97
The frequencies of infestations by the desert locust in different territories, <i>Z. Waloff and J.M. Connors</i>	105
Recommended analytical methods for pesticides	
27. Technical BHC dusts and water-dispersible powders	
<i>Collaborative Pesticides Analytical Committee</i>	114
Outbreaks and new records	
India	117
Plant quarantine announcements	
Honduras, Sweden	118
News and notes	120

No. 6 - DECEMBER

World citrus problems	
IV. Venezuela, <i>L.C. Knorr, G. Malaguti, D. Serpa and F. Leal</i>	121
Control of algal weeds in paddy fields in West Bengal, India, <i>S.K. Mukherji and S.K. Sengupta</i>	129
Recommended analytical methods for pesticides	
28. Technical DDT solutions, emulsifiable concentrates and emulsions	
<i>Collaborative Pesticides Analytical Committee</i>	131
Outbreaks and new records	
Israel, Togo, Venezuela	139
Plant quarantine announcements	
France, Netherlands	142
News and notes	144

INDEX

Acacia arabica var. *nilotica*, pest, Egypt, 139
Aceria sheldoni, citrus, Syria, 40, 41
Achatina fulica, Hawaii, 36
Acrobasis caryaef, pecan nut, United States, 28
Acromyrmex octospinosus, citrus, Venezuela, 127
Acrosternum hilare, soybean, United States, 27
Aculus schlechtendali, pear, United States, 28
Aculopsiphon pisum, alfalfa, United States, 27; pea, United States, 30
Aedes sollicitans, cattle, United States, 31
Aedes vexans, United States, 31
Agallia albida, citrus, Brazil, 101
Agallia ensigera, citrus, Brazil, 101
Agaliopsis novella, citrus, Brazil, 101
Ageneotettix deorum, United States, 25
Agrotis ipsilon, United States, 93
Albizia lebbeck, pest, Egypt, 139
Aleurocanthus hansfordi, citrus, Rhodesia, 22
Alfalfa, nematodes, United Arab Republic, 55; pests, United States, 26, 27
Almond, nematodes, United Arab Republic, 51, 55
Alsophila petrometaria, hickory, oak, ornamentals, United States, 34
Alternaria citri, citrus, Venezuela, 125
Alternaria sesami, sesame, 13
Alternaria sp., coconut, Togo, 141
Alternaria tenuis, sesame, 13
Alternaria triticina, wheat, India, 117
Amarantus sp., rice, India, 129
Amblyomma americanum, United States, 32
Ambracius dufourei, predator of *Orthezia praelonga*, Brazil, 92
Amphimallon majalis, United States, 36
Amphitormus coloradus, United States, 25
Angeles, Néstor de J., 141
Anthonomus grandis, cotton, United States, 30
Anticarsia gemmatalis, soybean, United States, 27
Antonina natalensis, grass, Rhodesia and Nyasaland, 22
Antonidiella aurantiii, citrus, Syria, 40
Apanteles artonae, hyperparasite on *Artona catoxantha*, Federation of Malaya, 42
Aphelechnoides sp., grass, sugarcane, United Arab Republic, 50, 54, 55
Aphis craccivora, citrus, Syria, 38, 40
Aphis gossypii, citrus, Syria, 38, 40; Venezuela, 126
Aphis spiraeae, citrus, Venezuela, 126
Aphrophora saratogensis, United States, 34
Aphytis chrysomphali, parasite on *Aspidiotus destructor*, Federation of Malaya, 42
Aphytis hispanicus, parasite on *Parlatoria pergandii*, Israel, 84
Apogonia cribicollis, coconut, Federation of Malaya, 43
Apple, pests, United States, 27, 28
Arepuna sp., citrus, Brazil, 101
Aroga websteri, sagebrush, United States, 27
Artona catoxantha, coconut, Federation of Malaya, 42
Ascochyta citri, citrus, Venezuela, 125
Aspidiophagus citrinus, parasite on *Aspidiotus destructor*, Federation of Malaya, 42
Aspidiophagus destructor, palms, Federation of Malaya, 42
Aspidiophagus perniciosus, peach, United States, 28
Atta sexdens, citrus, Venezuela, 127
Aulocara elliotti, United States, 25
Avocado, pests, United States, 35
Balsam, pests, United States, 34
Banana, diseases, Samoa, 79, 80; nematodes, United Arab Republic, 50, 51, 52, 55, 56
Barley, nematodes, United Arab Republic, 50, 55
Basiria graminophila, eggplant, United Arab Republic, 50
Bean, pests, United States, 29
Birch, pests, United States, 35
Blissus leucopterus, grass, United States, 27
Blueberry, pest, United States, 28
Bottiyodiplodia palmarum, coconut, Samoa, 80
Bottiyodiplodia sp., coconut, Togo, 140
Bottiyodiplodia theobromae, cocoa, Samoa, 80
Bougainvillea, pests, Israel, 82
Brachymeria euploae, hyperparasite on *Ptychomyia remota*, Federation of Malaya, 42
Brachyrhinus ovatus, United States, 32
Breadfruit, disease, Samoa, 81
Brevicoryne brassicae, cabbage, United States, 29
Brevipalpus obovatus, citrus, Syria, 41; Venezuela, 122, 125
Brevipalpus phoenicis, citrus, Brazil, 100; Venezuela, 122, 123, 125, 126
Brevipalpus sp., Syria, 41
Brychus pisorum, pea, United States, 30
Bryobia praetiosa, United States, 32
Buphonella murina, maize, Rhodesia, 22
Calonectria rigidiuscula, cocoa, Samoa, 80
Camnula pellucida, United States, 25
Camponotus sp., United States, 32
Capnodium sp., citrus, Brazil, 92
Carpocapsa pomonella, apple, United States, 27
Carter, Walter, 67, 73
Casca parvipennis, parasite on *Aspidiotus destructor*, Federation of Malaya, 42
Cassava, disease, Samoa, 81
Cephaleuros sp., citrus, Venezuela, 124
Ceraphron sp., parasite on *Artona catoxantha*, Federation of Malaya, 42
Ceratitidis capitata, United States, 28
Cercospora coffeicola, coffee, Samoa, 80
Cercospora henningsii, cassava, Samoa, 81
Cercospora kopkei, sugarcane, Samoa, 81
Cercospora musae, banana, Samoa, 79
Cercospora sesami, sesame, 13
Cercospora sesamicola, sesame, 13
Cercospora traversiana, fenugreek, 14
Cereals, pests, United States, 25
Ceroplastes floridensis, citrus, Syria, 40
Cerotoma trifurcata, soybean, United States, 27
Chalcoscelis albifasciata, coconut, Federation of Malaya, 43
Chara sp., rice, India, 129
Cheletogenes ornatus, predator on scale insects, Israel, 85
Cheletomimus berlesei, predator on *Hemiberlesia lata*, United States, 85
Cheletomimus sp., predator on *Parlatoria* sp., Israel, 85
Chermes piceae, balsam, United States, 34
Chilocorus bipustulatus, predator on *Parlatoria* sp., Israel, 84
Chilocorus stigma, 85
Choristoneura fumiferana, spruce, United States, 33
Chorizagratis auxiliaris, wheat, United States, 25
Chrysomphalus aonidum, citrus, Venezuela, 126

Chrysomphalus ficus, citrus, Syria, 40
Chrysops sp., United States, 32
Cicadulina mbila, virus vector, maize, Rhodesia and Nyasaland, 22
Cicadulina parazeae, virus vector, maize, Rhodesia and Nyasaland, 22
Cicadulina storeyi, virus vector, maize, Rhodesia and Nyasaland, 22
Circulifer tenellus, virus vector, sugar beet, United States, 30
 Citrus, blackfly, Rhodesia and Nyasaland, 22; diseases, Philippines, 1; Samoa, 81; Syria, 37-41; Venezuela, 121-128; false exanthema, Brazil, 94-104; mites, United States, 28, 29; nematodes, United Arab Republic, 50-52, 54-56; pests, Brazil, 92; Israel, 82; Syria, 37-41; United States, 29; viruses, 57-66.
Citrus aurantium, pests, Israel, 83
Citrus limetta, pests, Israel, 83
Citrus limon, pests, Israel, 83
Citrus macrophylla, 59
Citrus moi, 59
Citrus paradisi, pests, Israel, 83
Citrus reticulata, pests, Israel, 83
Citrus sinensis, pests, Israel, 83
Citrus volkameriana, 59
Cladosporium herbarum var. *citricolum*, citrus, Venezuela, 122
Cladosporium sp., coconut, Togo, 141
Clasterosporium coccicola, coconut, Togo, 141
 Clover, pests, United States, 27
Coccus viridis, citrus, Venezuela, 126
Cochliomyia hominivorax, United States, 32
 Cocoa, diseases, Samoa, 80
 Coconut, diseases, Samoa, 80; Togo, 140, 141; lethal yellowing, Jamaica, 67, 73-78; pests, Federation of Malaya, 42, 43; Samoa, 80; pests' parasites and predators, Federation of Malaya, 42, 43
Coccus nucifera, diseases, Jamaica, 73; Togo, 140
 Coffee, diseases, Samoa, 80
Coleophora laricella, larch, United States, 34
Colias eurytheme, alfalfa, United States, 27
Colletotrichum gloeosporioides, citrus, Venezuela, 124
Colletotrichum sp., coconut, Togo, 141
Coloradria pandora, Jeffrey pine, United States, 33, 34
 Commonwealth Entomological and Plant Pathological Conferences, 48
 Congress for Plant Protection (Second World), 120
 Congress on Tropical Plant Protection, 120
 Conners, J.M., 105
Conoderus falli, tobacco, United States, 30
Conoderus vespertinus, tobacco, United States, 30
Conotrachelus nenuphar, plum, United States, 28
Contarinia sorghicola, sorghum, United States, 26
Cordana musae, banana, Samoa, 80
Corticium salmonicolor, cocoa, Samoa, 80
Corynespora cassiicola, papaya, Samoa, 81
Corynespora sp., sesame, 13
 Costa, A.S., 98
 Cotton, nematodes, United Arab Republic, 50-56; pests, Rhodesia, 22; United States, 30, 31
 Cottonwood, pests, United States, 35
Crambus sp., grass, United States, 27
Cratosomus punctulatus, citrus, Venezuela, 126
Criconemoidea mutabile, grass, United Arab Republic, 50
Criconemoidea sp., citrus, United Arab Republic, 50, 54-56
 Crucifers, nematodes, United Arab Republic, 50, 51, 55; pests, United States, 29
Cryptococcus fagi, United States, 34
Crysomphalus dictyospermi, citrus, Syria, 40
Ctenocephalides sp., United States, 32
 Current, pest, United States, 28
Cybocephalus semiflavus, predator on scale insects, Federation of Malaya, 42
Cybocephalus sp., predator on scale insects, Federation of Malaya, 42
Cylindrosporium sesami, sesame, 13
Cynodon sp., rice, India, 129
Cyperus sp., rice, India, 129
 Date palm, nematodes, United Arab Republic, 55
Dendroctonus brevicomis, western pine, United States, 33
Dendroctonus engelmanni, spruce, United States, 33
Dendroctonus frontalis, United States, 33
Dendroctonus jeffreyi, Jeffrey pine, United States, 33
Dendroctonus monticolae, mountain pine, United States, 33
Dendroctonus ponderosae, United States, 33
Dendroctonus terebrans, United States, 33
Dereodus recticollis, maize, cotton, Rhodesia, 22
Dermacentor nitens, United States, 32
Dermestes lararius, United States, 32
Diabrotica longicornis, maize, United States, 26
Diabrotica undecimpunctata howardi, groundnuts, United States, 27
Diabrotica virgifera, maize, United States, 26
Diaphorina citri, citrus, Brazil, 101
Diaporthe citri, citrus, Venezuela, 124
Dictyothrinium quadratum, coconut, Togo, 141
Drimococcia javanica, parasite on *Promecotheca cumingi*, Federation of Malaya, 43
Diplodia natalensis, citrus, Venezuela, 125
Diplodia sp., coconut, Togo, 141
Ditylenchus sp., barley, grass, squash, United Arab Republic, 50, 54, 55
Drepanoptera femoratum, United States, 25
Echinocloa sp., rice, India, 129
 Eggplant, nematodes, United Arab Republic, 50, 55
Eichornia crassipes, Senegal, 93
Elasmopalpus lignosellus, groundnuts, United States, 27
 Elm, pests, United States, 34, 35
Elsinoe australis, citrus, Venezuela, 123
Elsinoe fawcettii, citrus, 60; Samoa, 81; Venezuela, 123
Emapoasca fabae, alfalfa, United States, 27; potato, United States, 29
Emapoasca krameri, citrus, Brazil, 101
Ennomos subsignarius, elm, United States, 34
Epicoccum coccum, coconut, Samoa, 80
Epilachna varivestis, bean, United States, 29; soybean, United States, 27
Epitrimerus pyri, pear, United States, 28
Epitriz cucumeris, tomato, United States, 29
Epitriz hirtipennis, tobacco, United States, 30
Eryphyes pyri, pear, United States, 28
Estigmene acrea, cotton, United States, 31
Euplemone catoxanthae, parasite on *Parasa lepida*, Federation of Malaya, 43
 European pine, pests, United States, 34
Eurytoma albobilialis, parasite on *Artona catoxantha*, Federation of Malaya, 42
Eurytoma sp., parasite on *Parasa lepida*, Federation of Malaya, 43

Euschistus servus, soybean, United States, 27
Eutetranychus bankei, citrus, United States, 29; Venezuela, 126
Exosporium sp., coconut, Togo, 141

FAO/IUFRO Symposium on Forest Diseases and Insects, 48
Feltia subterranea, soybean, United States, 27
Fenugreek, disease, 14
Fenusia pusilla, birch, United States, 35
Ficus pseudosicomora, pest, Egypt, 139
Fig, nematodes, United Arab Republic, 51, 52, 55; pest, Egypt, Israel, 139
Fir, pests, United States, 33, 34
Forage, pests, United States, 25
Forficula auricularia, United States, 32
Frankliniella vaccinii, blueberry, United States, 28
Fusarium oxysporum f. *cubense*, banana, Samoa, 80
Fusarium sp., coconut, Togo, 140; on *Orthezia praelonga*, Brazil, 93

Galerucella xanthomelaena, elm, United States, 35
Gitona brasiliensis, predator of *Orthezia praelonga*, Brazil, 92
Gloeosporium limetticolum, citrus, Venezuela, 124
Gloeosporium sp., coconut, Togo, 140
Glomerella cingulata, coffee, Samoa, 80
Glomerella tucumanensis, sugarcane, Samoa, 81
Gnorimoschema operculella, potato, United States, 29; stored products, United States, 35
Concalves, Cincinato R., 92
Goplana dioscoreae, yam, Samoa, 81
Goriphius sp., parasite on *Chalcocelis alboguttata*, Federation of Malaya, 43
Graminella colones, citrus, Brazil, 101
Grapevine, nematodes, United Arab Republic, 49, 50, 51, 52, 55; pests, Egypt, 139; Israel, 139; United States, 35
Graphognathus sp., United States, 36
Grasses, nematodes, United Arab Republic, 50, 51, 55, 56; pests, Rhodesia and Nyasaland, 22; United States, 27
Groundnuts, nematodes, United Arab Republic, 55; pests, United States, 27
Guava, nematodes, United Arab Republic, 55
Gymnandrosoma aurantianum, citrus, Venezuela, 126

Haeceliania brontispa, parasite on *Plesispa reichei*, *Brontispa mariana*, Federation of Malaya, 42, 43
Haematochia irritans, United States, 31
Helicotylenchus costenbrinki, grape, grass, United Arab Republic, 50
Helicotylenchus dihystera, cabbage, citrus, eggplant, grape, grass, olive, squash, sugarcane, United Arab Republic, 49, 50
Helicotylenchus egyptiensis, sugarcane, United Arab Republic, 50
Helicotylenchus microlobus, cotton, United Arab Republic, 50
Helicotylenchus multicinctus, citrus, grape, grass, United Arab Republic, 49, 50
Helicotylenchus similis, citrus, United Arab Republic, 50
Heliothis sp., tobacco, United States, 30
Heliothis virescens, cotton, United States, 30; tobacco, United States, 30
Heliothis zea, cotton, United States, 30; lettuce, United States, 29; maize, United States, 26; sorghum, United States, 26; soybean, United States, 27; tomato, United States, 29
Helminthosporium sp., coconut, Togo, 140
Hemerocampa pseudotsugata, Douglas-fir, United States, 34
Hemberlesia lataniae, United States, 84
Hemicyclophora sp., United Arab Republic, 54, 56; citrus, grape, grass, United Arab Republic, 55
Hemileia vastatrix, coffee, Samoa, 80
Hemileuca olivea, United States, 27
Hemisarcopes coccophagus, predator on *Parlatoria* sp., Israel, 84; on *Aonidiella aurantii*, South Africa, 84
Hemisarcopes malus, predator of scale insects, 84
Hemitarsonemus latus, citrus, Brazil, 101
Hemlock, pests, United States, 34
Hickory, pests, United States, 28
Hirschmanniella oryzae, cotton, United Arab Republic, 50
Homadaula albizziae, mimosa, honeylocust, United States, 35
Honeylocust, United States, 35
Hoplalinus columbus, banana, citrus, cotton, sugarcane, United Arab Republic, 49, 50
Hylemya brassicae, cabbage, United States, 29
Hylemya floralis, crucifers, United States, 29
Hypera postica, alfalfa, United States, 26, 93
Hyphantria cunea, United States, 34
Hypoderma sp., United States, 32

Icerya purchasi, citrus, Venezuela, 126
International Meeting on Rice Production and Protection, 24

Janus integer, currant, United States, 28
Jeffrey pine, pests, United States, 33, 34
Johnston, A., 79
Jujube, nematodes, United Arab Republic, 50, 52, 55

Knorr, L.C., 37, 121

Longidorus taniwha, citrus, fig, grass, United Arab Republic, 51
Loxostege sticticalis, sugar beet, United States, 30
Lugus sp., cotton, United States, 31

Macraspis lucida, maize, Venezuela, 141
Macrodactylus subspinosus, ornamentals, United States, 35
Macrosiphum avenae, small grains, United States, 25
Maize, nematodes, United Arab Republic, 51, 52, 55; pests, Rhodesia and Nyasaland, 22; United States, 26, 27, 28; Venezuela, 141; virus, Rhodesia and Nyasaland, 22
Malacosoma americanum, ornamentals, United States, 34
Malacosoma disstria, oak, sweetgum, tupelogum, United States, 35
Malacosoma fragile, cottonwoods, United States, 35
Malacosoma pluviale, ornamentals, United States, 35
Malaguti, G., 121
Mango, disease, Samoa, 81; nematodes, United Arab Republic, 55
Manser, P.D., 140
Marasmiellus sp., banana, Samoa, 80
Margaderes sp., sugarcane, Rhodesia and Nyasaland, 22
Martinez, A.L., 1
Matsucoccus resinosae, red pine, United States, 34

Meeting on Coconut Production and Protection, FAO, 144

Melanoplus bivittatus, United States, 25

Melanoplus differentialis, United States, 25

Melanoplus femur-rubrum, United States, 25

Melanoplus packardi, United States, 25

Melanoplus sanguinipes, United States, 25

Melititobia sp., hyperparasite on *Ptychomyia remota*, Federation of Malaya, 42

Meloidogyne javanica, fig, United Arab Republic, 51, 54

Mesoleuros dentipes, maize, tobacco, Rhodesia, 22

Microlarinus lareynii, United States, 36

Microlarinus lypriformis, United States, 36

Microrotidea sp., hyperparasite on *Artona catoxantha*, Federation of Malaya, 42

Mimosa, pest, United States, 35

Moreira, S., 57

Mountain pine, pests, United States, 33

Mukherji, S.K., 129

Mulberry, nematodes, United Arab Republic, 52, 55

Müller, Gerd W., 98

Musca autumnalis, United States, 31

Musca domestica, United States, 31

Mycosphaerella coffeeae, coffee, Samoa, 80

Mycosphaerella musicola, banana, Samoa, 79

Myzus persicae, citrus, Syria, 38, 40; Venezuela, 126; spinach, United States, 30; tobacco, United States, 30

Nectria cinereo-papillata, cocoa, Samoa, 80

Nectria ochroleuca, cocoa, Samoa, 80

Nectria sp., United States, 34

Neodiprion abietis, white fir, United States, 34

Neodiprion sertifer, European pine, 34

Neotylenchus sp., almond, United Arab Republic, 51, 54, 55

Nezara viridula, soybean, United States, 27

Nezara viridula var. *smaragdula*, ornamentals, vegetables, United States, 30

Nipa sp., pest, Federation of Malaya, 42

Nitella sp., rice, India, 129

Oak, pests, United States, 34, 35

Oedaleonotus enigma, United States, 25

Oligonychus pratensis, maize, United States, 27

Olive, nematodes, United Arab Republic, 50, 55; pest, Israel, 139

Oenocrytus podontiae, parasite on *Plesispa reichei*, Federation of Malaya, 42

Ornamentals, pests, United States, 30, 34, 35

Orthzia praelonga, citrus, Brazil, 92

Oryctes rhinoceros, coconut, Samoa, 80

Oryzaephilus surinamensis, stored products, United States, 35

Ostrinia nubilalis, maize, United States, 26

Oulema melanopa, cereal, United States, 25

Paleacrita vernata, United States, 34

Palms, pests, Federation of Malaya, 42

Panicum sp., rice, India, 129

Panonychus ulmi, apple, United States, 27, 28

Papaya, disease, Samoa, 81

Papilio cresphontes, citrus, United States, 29

Parasa lepida, coconut, Malaya, 43

Paratriozia cockerelli, potato, United States, 29

Parlatoria cinerea, citrus, Israel, 82

Parlatoria oleae, Israel, 84

Parlatoria pergandii, citrus, Israel, 82; Syria, 40

Paropota paradoxus, fig, grapevine, olive, Israel, 139

Paspalum sp., rice, India, 129

Pea, pests, United States, 30

Peach, nematodes, United Arab Republic, 55; pests, United States, 28

Pear, moria, Italy, 6, 11; nematodes, United Arab Republic, 51, 52, 54, 55; pests, United States, 28

Pecan nut, pests, United States, 28

Pectinophora gossypiella, cotton, United States, 31

Pediobius detritus, hyperparasite on *Ptychomyia remota*, Federation of Malaya, 42

Pediobius parvulus, hyperparasite on *Ptychomyia remota*, Federation of Malaya, 42

Pediobius sp. (= *Pleurotropis* sp.), parasite on *Parasa lepida*, Federation of Malaya, 43

Pelidnota laevissima, maize, Venezuela, 141

Pemphigus populivorus, cabbage, United States, 29

Pentalonia nigronervosa, banana, Samoa, 79

Pepper, pests, United States, 29

Perilampus sp., parasite on *Xylotachina* sp., Federation of Malaya, 43

Pestalotia palmarum, coconut, Togo, 140, 141

Pestalotiopsis palmarum, coconut, Samoa, 80

Petrobia latens, wheat, United States, 26

Phoma sp., coconut, Togo, 141

Phyllocoptrus oleivora, citrus, Syria, 41; United States, 28; Venezuela, 125

Phyllostictina musarum, banana, Samoa, 80

Phytophaga destructor, wheat, United States, 26

Phytophthora palmivora, cocoa, Samoa, 80

Phytophthora parasitica, citrus, Venezuela, 122

Phytophthora sp., citrus, Brazil, 59, 60; Venezuela, 121, 122

Pieris rapae, cabbage, cauliflower, United States, 29

Pine, nematodes, United Arab Republic, 55

Pissodes strobi, white-pine, United States, 33

Planoecoccus citri, citrus, Syria, 40

Plant Quarantine Conference in the South Pacific (Second Regional), 24

Plasmopara halstedii, sunflower, 14

Plathypena scabra, soybean, United States, 27

Platytius bicolor, Brazil, 97, 100-103

Plaut, H.N., 139

Plesispa reichei, coconut, *Nipa*, royal palm, Federation of Malaya, 42

Pleurotropis parvulus, (= *Pediobius*), parasite on *Pro-mecotheca cumingi*, Federation of Malaya, 43

Plodia interpunctella, stored products, United States, 35

Plum, insects, United States, 28

Poncirus trifoliata, 57, 59, 60, 62; Brazil, 98

Popillia japonica, United States, 36

Porthetria dispar, United States, 34

Potato, pests, United States, 29

Pratylenchoides sp., pear, citrus, United Arab Republic, 51, 54, 55

Pratylenchus crenatus, sugarcane, United Arab Republic, 51

Pratylenchus thornei, cabbage, cotton, fig, maize, sugarcane, United Arab Republic, 51, 54

Pratylenchus zeae, citrus, United Arab Republic, 51

Promecotheca cumingi, coconut, Federation of Malaya, 43

Prosapia bicincta, grasses, United States, 27

Prospaltella inquirenda, parasite on *Parlatoria oleae*, *Parlatoria pergandii*, Israel, 84

Protoparce sp., tobacco, United States, 30

Pseudaleitia unipuncta, United States, 93; small grains, United States, 25

Pseudaulacaspis pentagona, peach, United States, 28

Pseudhalenchus anchilisposomus, pear, United Arab Republic, 51

Pseudhalenchus sp., grass, United Arab Republic, 51, 54, 55, 56

Pseudococcus sp., citrus, Venezuela, 126

Psilenchus aestuarius, cotton, United Arab Republic, 51

Psilenchus hilarius, cotton, United Arab Republic, 51

Psilenchus magnidens, cotton, fig, grass, United Arab Republic, 51

Psoroptes ovis, United States, 32

Psylla pyricola, pear, Italy, 11; United States, 28

Psylla pyris, pear, Italy, 11

Ptychomyia remota, parasite on *Artona catoxantha*, Federation of Malaya, 42

Puccinia kuehnii, sugarcane, Samoa, 81

Pyrus communis, see pear

Pyrus serotina, see pear

Pyrus ussuriensis, see pear

Quarantine legislation, Australia, 23; France, 142; Honduras, 94-96, 118-119; Ivory Coast, 23; Luxembourg, 23; Netherlands, 143; United Kingdom, 46, 72; Upper Volta, 72

Radopholus similis, citrus, Venezuela, 124; pear, United Arab Republic, 51, 54

Red pine, pests, United States, 34

Refatti, E., Italy, 6

Rhagoletis completa, walnut, United States, 28

Rhagoletis pomonella, apple, United States, 28

Rhipicephalus sanguineus, United States, 32

Rhopalosiphum fitchii, small grains, United States, 25

Rhopalosiphum maidis, maize, United States, 26; small grains, United States, 25

Rhyacionia buoliana, European pine, United States, 34

Rice, blast, Southeast Asia, 24; nematodes, pests, 24, 25; viruses, Southeast Asia, 24; weeds, India, 129

Rotylenchulus reniformis, citrus, grape, sugarcane, United Arab Republic, 51

Rotylenchulus sp., United Arab Republic, 54; almond, citrus, eggplant, United Arab Republic, 55

Sagebrush, pests, United States, 27

Saissetia coffeeae, citrus, Syria, 40

Saissetia hemisphaerica, citrus, Venezuela, 126

Saissetia nigra, citrus, Venezuela, 126

Sannioidea exitiosa, peach, United States, 28

Sarcophaga antilope, parasite on *Parasa lepida*, Federation of Malaya, 43

Schistocerca gregaria, 105

Schizaphis graminum, small grains, United States, 25

Scolecosidium sp., taro, Samoa, 81

Scolytus ventralis, fir, United States, 33

Scutellonema brachyurum, grape, grass, United Arab Republic, 51

Scymus sp., predator of *Orthezia paelonga*, Brazil, 92

Seal, William L., 25

Selanaspis articulatus, citrus, Venezuela, 126

Sengupta, S.K., 129

Septobasidium sp., citrus, Venezuela, 124

Septoria citri, citrus, Venezuela, 125

Serpia, D., 121

Sesame, diseases, 13

Sesamum indicum, see Sesame

Sesamum orientale, see Sesame

Seynesia sp., coconut, Togo, 1

Simulium sp., United States, 31

Siphoninus finitimus, citrus, Syria, 40

Small grains, pests, United States, 25, 26

Sogata orizicola, rice, United States, 25

Solenopsis saevissima richteri, United States, 36

Sorghum, pests, United States, 26

Sowell Jr., G., 13

Soybean, pests, United States, 27

Spaniopterus crucifer, parasite on *Aspidiota destructor*, Federation of Malaya, 42

Sphaceloma fawcetti, var. *scabiosa*, citrus, Venezuela, 123

Sphaeropsis tumefaciens, citrus, Venezuela, 125

Spinach, pests, United States, 30

Spissistilus festinus, alfalfa, United States, 27

Spodoptera exigua, cotton, United States, 31; soybean, United States, 27

Spruce, pests, United States, 33

Squash, nematodes, United Arab Republic, 50, 55

Stigmella mangiferae, mango, Samoa, 81

Stomoxys calcitrans, United States, 32

Stored products, pests, United States, 35

Sugar beet, pests, United States, 30

Sugarcane, diseases, Samoa, 81; nematodes, United Arab Republic, 50, 51, 52, 54, 55, 56; pests, Rhodesia and Nyasaland, 22

Sunflower, disease, 14

Supella supellecillum, United States, 32

Sweetgum, pest, United States, 35

Synanthedon pictipes, peach, United States, 28

Syntomosphyrum obscuripes, hyperparasite on *Ptycho-myia remota*, Federation of Malaya, 42

Systates exaptus, maize, Rhodesia, 22

Tabanus sp., United States, 32

Tarjan, A.C., 49

Taro, disease, Samoa, 81

Telenomus basalis, parasite on *Nezara viridula*, var. *smaragdula*, United States, 30

Telytolenchus ventralis, grass, United Arab Republic, 51

Tetranychus atlanticus, clover, United States, 27; cotton, United States, 31; soybean, United States, 27

Tetranychus cinnabarinus, cotton, United States, 31

Tetranychus medanieli, apple, United States, 28

Tetranychus telarius, apple, United States, 27; clover, United States, 27; cotton, United States, 31

Tetrastichus brontispae, parasite on *Brontispa mariana*, Federation of Malaya, 43

Theroaphis maculata, alfalfa, United States, 27

Thylenchorhynchus clarus, maize, United Arab Republic, 52

Thylenchorhynchus clavicaudatus, citrus, United Arab Republic, 52

Thylenchorhynchus kegenicus, cotton, United Arab Republic, 52

Thylenchorhynchus latus, citrus, cotton, fig, jujube, pear, United Arab Republic, 49, 52

Thylenchorhynchus nothus, citrus, eggplant, maize, mulberry, United Arab Republic, 52

Thylenchulus semipenetrans, citrus, United Arab Republic, 52, 54, 56; grape, United Arab Republic, 52

Thylenchus sp., pear, United Arab Republic, 52, 54, 55

Thyridopterix ephemeraeformis, ornamentals, United States, 35

Tipha sp., parasite on *Apogonia cribicollis*, Federation of Malaya, 43
 Tobacco, pests, Rhodesia, 22, United States, 30
 Tomato, pests, United States, 29
Toxoptera aurantii, citrus, Syria, 38, 40; Venezuela, 126
Toxoptera citricidus, citrus, Brazil, 101; Philippines, 3; Venezuela, 126; not found in citrus, Syria, 38, 40
Tribolium sp., stored products, United States, 35
Tribulus cistoides, Hawaii, 36
Tribulus terrestris, United States, 36
Trichodorus minor, grape, United Arab Republic, 51
Trichodorus teres, citrus, United Arab Republic, 51
Trichoplusia ni, cotton, United States, 31; crucifers, United States, 29; soybean, United States, 27; tobacco, United States, 30
Trichopoda pennipes, parasite on *Nezara viridula* var. *smaragdula*, United States, 30
Trigona amalthea, citrus, Venezuela, 127
Trigona trinidadensis, citrus, Venezuela, 127
Trogoderma granarium, stored products, United States, 35
Tupelogum, pest, United States, 35
Tylenchorhynchus brassicae, citrus, grass, pear, United Arab Republic, 51

Unaspis citri, citrus, Venezuela, 126
Uredo artocarpi, breadfruit, Samoa, 81
Uromyces musae, banana, Samoa, 80

 Vasudeva, R.S., 140

 Vaughn, J.R., 37
Verticillium lecanii, on *Orthezia praelonga*, Brazil, 93

 Wallace, J.M., 1
 Walnut, pests, United States, 28
 Waloff, Z., 105
 Water hyacinth, Senegal, 93
 Western pine, pests, United States, 33
 Wheat, pests, United States, 25, 26
 Whellan, J.A., 22
 White pine, pests, United States, 33

Xiphinema americanum, grape, United Arab Republic, 52
Xiphinema arenarium, citrus, fig, United Arab Republic, 52
Xiphinema elongatum, citrus, United Arab Republic, 52
Xiphinema insigne, banana, citrus, grape, sugarcane, United Arab Republic, 52
Xiphinema sp., mulberry, United Arab Republic, 52, 54, 55, 56
Xylosandrus compactus, avocado, grape, United States, 35
Xylotachina sp., parasite on *Apogonia cribicollis*, Federation of Malaya, 43

 Yam, disease, Samoa, 81

 Zakari, I., 140
Zeadiaatraea grandiosella, maize, United States, 26
Zonostemata electa, pepper, United States, 29

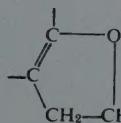
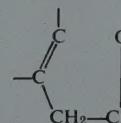
ERRATA

(The line numbers do not take account of titles and headings)

Page 30, column 2, line 18 *read Epitrix for Epitix*
 " 71, " 2, footnote 4 " acetate " arsenate
 " 80, " 1, line 17 " *oxysporum* " *oxyporum*
 " 92, name of author " Gonçalves " Gonçales

ERRATA AND ADDITIONS TO BE INCLUDED IN ISSUES OF THE BULLETIN
PRIOR TO 1964

(The line numbers do not take account of titles and headings)

Volume 9, page 19, column 2, Table I	read 3(i), 3(ii), 3(iii), 3(iv)	for 3(i), 3(ii), 3(iii), 3(iv)
" 9, " 27, " 1, between lines 20 and 21 insert	Heat under reflux until the mixture is boiling gently, add the condenser zinc wool (1 gm) rolled into a ball, continue boiling for 15 minutes, repeat with a further two lots of zinc wool (3 x 1 gm in all). Remove the source of heat, rinse the condenser with a little acetic acid, allow the apparatus to cool until the flask can be conveniently handled, and detach the flask (Note 2). Fit the Buchner funnel with two 5.5-cm filter papers and cover them with a layer of Gooch asbestos (Notes 1 and 3). Decant the hot contents of the flask on to the filter, taking care to retain the zinc, which may have partly decomposed, in the flask as far as possible. Rinse the flask with water and pass the rinsings through the filter. Wash the residue in the flask and on the filter with water until the washings are halide-free, i.e., give no precipitate on testing with silver nitrate solution.	
" 9, " 154, " 1, line 38	read Determine the chloride ion either for by the Volhard method using the thiocyanate or electrometrically.	Titrate the solution with the thiocyanate or determine the chloride ion electrometrically.
" 10, " 67, " 1, structural formula	"	"
" 10, " 112, " 1, line 34		
" 10, " 112, " 2, " 40	" w/w = $\frac{304.45 \times t}{w}$	" w/w = $\frac{304.45 - t}{w \times 100}$
" 10, " 113, " 1, " 46	" "	" "
" 10, " 114, " 1, " 8	" "	" "
" 10, " 134, " 1, " 2	" 290.9	" 296.5
" 10, " 135, " 2, " 8	" Determine the chloride ion by the Volhard method using the thiocyanate solution or electrometrically.	Titrate the solution with the thiocyanate or determine the chloride ion electrometrically.

Volume 11, page 110, column 1, structural formula

read NO₂

for N₂O

" 11, " 111, " 1, after line
10 insert

Determine the melting point of the isolated DNOC by the prescribed method MT 2 (below). This should not be less than 84°C and should not be depressed by admixture with an equal quantity of pure DNOC.

" 11, " 112, " 1, line 2

read 255.5

" 176.9

" 11, " 132, " 2, " 10

" (27 mg)

" (27 gm)

